

## Generation of $\alpha$ -Nitroalkyl Radicals by Oxidation of Nitronate Anions with Cerium(IV) Ammonium Nitrate and Their Addition Reaction to Electron-Rich Olefins

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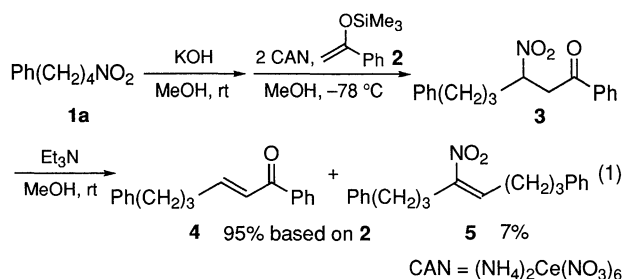
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$\alpha$ -Nitroalkyl radicals are generated by oxidation of nitronate anions with cerium(IV) ammonium nitrate. When the reactions are carried out in the presence of electron-rich olefins, such as silyl enol ethers, intermolecular addition of the radicals proceeds to afford  $\beta$ -nitroketones, which are further converted to  $\alpha,\beta$ -unsaturated ketones in good yield.

Though several attempts have been carried out to form carbon-carbon bonds by the use of carbon-centered radicals generated by the one-electron oxidation of carbanions, most of them afford self-coupling products.<sup>1-3</sup> Few examples have been developed for the intermolecular carbon-carbon bond forming reactions.<sup>4-6</sup> In these intermolecular reactions, the yield of the products is not sufficiently high, even with the use of large excess amounts of radical acceptors. We have been investigating the generation of carbon-centered reactive species by oxidation of  $\alpha$ -stannyl sulfides,<sup>7</sup>  $N$ -( $\alpha$ -stannylalkyl)amides,<sup>8</sup> or  $\alpha$ -stannyl esters<sup>9</sup> with cerium(IV) ammonium nitrate (CAN). In the course of this study,  $\alpha$ -nitroalkyl radicals were found to be generated by oxidation of nitronate anions with CAN, and intermolecular addition reaction proceeded with electron-rich olefins in high yield.<sup>10, 11</sup>

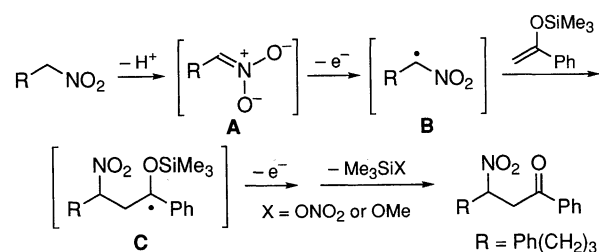
When a solution of potassium salt of 1-nitro-4-phenylbutane (**1a**) in methanol, prepared by treating **1a** with KOH, was added to a solution of CAN and  $\alpha$ -trimethylsilyloxystyrene (**2**) in methanol at  $-78^\circ\text{C}$  (**1a** : KOH : CAN : **2** = 1 : 1.3 : 1.9 : 0.85),<sup>12</sup> the reaction proceeded smoothly to give 3-nitro-1,6-diphenylhexan-1-one (**3**) as a crude product. Since purification of the nitroketone **3** by preparative TLC (silica gel) turned out difficult because of the elimination of nitrous acid from **3**, the crude product was purified after converting to 1,6-diphenyl-2-hexen-1-one (**4**). Thus, by treatment of the crude nitroketone **3** with triethylamine, the  $\alpha,\beta$ -unsaturated ketone **4** was obtained in 95% yield based on the silyl enol ether **2**, accompanied by the 5% recovery of **1a** and the formation of self-coupling product **5** (Eq. 1).



Even when the reaction was carried out by using exactly stoichiometric amounts of the reagents (**1a** : KOH : CAN : **2** = 1 : 1 : 2 : 1), the addition product **4** was obtained in 85% yield with the recovery of 13% of **1a**; the self-coupling product **5** was not

detected. It is well known that nitronate anions react with carbon electrophiles, such as alkyl halides, on the oxygen atom predominantly and selective  $C$ -alkylation is difficult.<sup>13, 14</sup> In the present radical reaction with the silyl enol ether, exclusive  $C$ -alkylation was observed as shown in Eq. 1.

The probable mechanism of the reaction is shown in Scheme 1. The nitronate anion **A** is oxidized to the  $\alpha$ -nitroalkyl radical **B**, which adds to the enol ether. Successive oxidation of the radical intermediate **C** and desilylation afford the product,  $\beta$ -nitroketone.



Scheme 1.

Use of other metallic oxidizing agents did not improve the yield of the product. The addition product **4** was obtained in 12% yield only when ferrocenium hexafluoro-phosphate was employed, and the use of  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ ,  $\text{Mn}(\text{pic})_3$ , or  $\text{Ag}(\text{pic})_2$  ( $\text{pic} = 2$ -pyridinecarboxylate) gave only the self-coupling product **5**, or brought about oxidation or hydrolysis of the silyl enol ether **2**. As for the base used for generating the nitronate, alkali metal hydroxides ( $\text{LiOH} \cdot \text{H}_2\text{O}$ ,  $\text{NaOH}$ ) and methoxides ( $\text{NaOMe}$ ,  $\text{KOMe}$ ) could be employed successfully, however, no reaction occurred by the use of triethylamine.

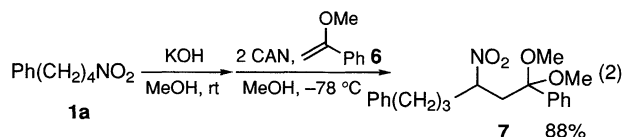
The reactions between various nitroalkanes and silyl enol ethers were examined and the results are summarized in Table 1. The nitroalkane **1a** reacted with several silyl enol ethers as well as the aryl silyl enol ether **2** to afford the corresponding addition products in high yield (Entries 1-3). When 5-nitropentyl acetate (**1b**) was employed, the addition products were obtained in moderate yield (Entries 4-6). A nitroalkane possessing two substituents at  $\alpha$ -position **1c** and a cyclic nitroalkane **1d** also afforded  $\alpha,\beta$ -unsaturated ketones in excellent yield in the reaction with **2** (Entries 7 and 9), however, the reactions with an aliphatic silyl enol ether resulted in poor yield (Entries 8 and 10). Phenylnitromethane (**1e**) gave a complex mixture and low yield of the product (Entry 11). In the case of nitromethane (**1f**), the oxidation did not proceed (Entry 12).<sup>15</sup> The configuration of the  $\alpha,\beta$ -unsaturated ketones obtained by these reactions was exclusively *E*.

Besides silyl enol ethers, vinyl ethers and styrene derivatives were available for the reaction. For instance,  $\alpha$ -methoxystyrene (**6**) reacted with the nitroalkane **1a**, giving a dimethylacetal of  $\beta$ -nitroketone **7** in high yield (Eq. 2).

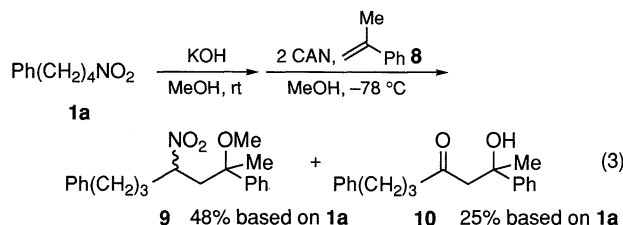
**Table 1.** The oxidative cross-coupling reaction between nitroalkanes and silyl enol ethers<sup>a</sup>

**1a** R<sup>1</sup> = Ph(CH<sub>2</sub>)<sub>3</sub>, R<sup>2</sup> = H  
**1b** R<sup>1</sup> = AcO(CH<sub>2</sub>)<sub>4</sub>, R<sup>2</sup> = H  
**1c** R<sup>1</sup>, R<sup>2</sup> = Me  
**1d** R<sup>1</sup>, R<sup>2</sup> = (CH<sub>2</sub>)<sub>5</sub>  
**1e** R<sup>1</sup> = Ph, R<sup>2</sup> = H  
**1f** R<sup>1</sup>, R<sup>2</sup> = H

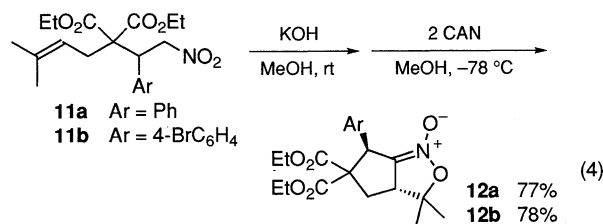
Entry	Nitroalkane	R <sup>3</sup>	R <sup>4</sup>	Yield / % <sup>b</sup>
1	<b>1a</b>	H	Ph	95
2	<b>1a</b>	H	<i>n</i> -Pr	81
3	<b>1a</b>	Me <sup>c</sup>	Ph	79
4	<b>1b</b>	H	Ph	69 <sup>d</sup>
5	<b>1b</b>	H	<i>n</i> -Pr	60 <sup>d</sup>
6	<b>1b</b>	Me <sup>c</sup>	Ph	65 <sup>d</sup>
7	<b>1c</b>	H	Ph	98 <sup>d</sup>
8	<b>1c</b>	H	Ph(CH <sub>2</sub> ) <sub>2</sub>	4 <sup>d</sup>
9	<b>1d</b>	H	Ph	90 <sup>d</sup>
10	<b>1d</b>	H	Ph(CH <sub>2</sub> ) <sub>2</sub>	43 <sup>d</sup>
11	<b>1e</b>	H	Ph	10
12	<b>1f</b>	H	Ph	0

<sup>a</sup> Reactions were carried out according to the procedure described in the text.<sup>b</sup> Based on the silyl enol ethers. <sup>c</sup> *E/Z* = 5/95 (determined by NMR).<sup>d</sup> The crude nitroketone was treated with DBU in CH<sub>2</sub>Cl<sub>2</sub> instead of triethylamine in methanol.

The reaction with  $\alpha$ -methylstyrene (**8**) afforded a methyl ether **9** (diastereomer mixture, 68 : 32) and a  $\beta$ -hydroxyketone **10**, though 4 molar amounts of **8** were required (Eq. 3).



The reaction can be applied for intramolecular cyclization. When nitro compounds **11a** and **11b** were treated with KOH, followed by oxidation with CAN, the cyclized products **12a** and **12b** were obtained as a single stereoisomer in good yield, respectively (Eq. 4). The relative configuration of **12b** was unambiguously determined as *trans* by X-ray crystallographic analysis.<sup>16</sup>



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- Though self-coupling products such as **5** were obtained in Entries 1-3 and 11, in other cases the corresponding dimers were not detected.
- Crystal data for **12b**: C<sub>20</sub>H<sub>24</sub>BrNO<sub>6</sub>, *M<sub>w</sub>* = 454.32, colorless prism, monoclinic, *P*2<sub>1</sub>/*a* (#14), *a* = 8.655(3), *b* = 19.595(4), *c* = 12.411(4) Å,  $\beta$  = 100.90(2)°, *V* = 2067.0(9) Å<sup>3</sup>, *Z* = 4, *D*<sub>calc</sub> = 1.460 g cm<sup>-3</sup>,  $\mu$  (Mo *K* $\alpha$ ) = 20.30 cm<sup>-1</sup>, *R* = 0.075, *R<sub>w</sub>* = 0.050, 3790 unique reflections.